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Introduction to Ionic Bonding and Covalent Bonding

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Section 7.3 – Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

Chapter 7 – Ionic and Metallic Bonding

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1. Chapter 7“Ionic and Metallic Bonding” 2. Section 7.1 - Ions OBJECTIVES: -Determine the number of valence electrons in an atom of a representative element. 3. Section 7.1 - Ions OBJECTIVES: -Explain how the octet rule applies to atoms of metallic and nonmetallic elements. 4.

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206 Chapter 7 □ Ionic Compounds and Metals Section 7.1.1 Figure 7.1 As carbon dioxide dis-solves in ocean water, carbonate ions are produced. Coral polyps capture these carbonate ions, producing crystals of calcium carbonate, which they secrete as an exo-skeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

Chapter 7: Ionic Compounds and Metals

Chapter 7 Ionic and Metallic Bonding Multiple Choice Identify the choice that best completes the statement or answers the question. ____ 1. What are the valence electrons of an atom? a. The electrons that an atom gains when it becomes an anion b. The electrons that remain closest to the nucleus c. The electrons in the highest occupied energy level d. The electrons that an atom loses when it ...

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chapter 7 notes chemistry bonding ionic metallic ...

Lesson 7.1 Reading and Study Workbook □ Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. 83. Ionic and Metallic Bonding. BONDING AND INTERACTIONS. 7.1 Ions. Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged. Lesson Summary. Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation. For a representative element, the group number equals the ...

BONDING AND INTERACTIONS

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Chemical Bonding Types of chemical bonds Lewis structures and the octet rule Ionic and covalent bonding Bond polarity and electronegativity Formal charge Chapter 8 Energetics of bond formation and dissociation Bond formation: $A + B \rightarrow AB$ If the reaction is energetically favorable, AB is lower in energy than $A + B$. Heat is released .

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