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CHAPTER 9 REVIEW Stoichiometry SECTION 3

PROBLEMS Write the answer on the line to the left.

Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) → 2NH<sub>3</sub>(g)

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through #4 on page 301 of the textbook. Section  
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with mass relationships in compounds and in chemical  
reactions.  $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$  2- a)  $2 \text{ mol HgO} \rightarrow 2 \text{ mol Hg} \text{ \& } 2 \text{ mol HgO} \rightarrow 1 \text{ mol O}_2$ ...

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